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AP Chemistry Solution Chemistry Practice Problems

Name _____ Date _____ Period _____

1. Consider a 1.00 L solution containing 0.10 M $\text{Al}(\text{OH})_3$, 0.10 M $\text{Fe}(\text{OH})_3$ and 0.10 M $\text{Ni}(\text{OH})_2$. What are the molar concentrations of the aluminum, iron and nickel ions?
2. In the net ionic equation for water in the net ionic equation for the reaction between NaOH and HNO_3 , explain.
3. Which of the following substances is a strong electrolyte in aqueous solution?
 NaCl , C_6H_6 , H_2O , $\text{C}_2\text{H}_5\text{OH}$
4. A 25.0 mL sample of an aqueous solution of NaOH requires 18.4 mL of 0.100 M HCl for its neutralization. What is the molarity of the NaOH solution?
5. Write net ionic equations for the following acid-base reactions:
 - a. The strong base KOH and the weak acid $\text{HC}_2\text{H}_3\text{O}_2$.
 - b. The strong base barium hydroxide and the strong acid phosphoric acid.
 - c. The weak base methyl amine (CH_3NH_2) and the strong acid hydrochloric acid (HCl).
6. What is the molarity of a solution of hydrochloric acid if exactly 10.0 mL of 0.100 M sodium hydroxide is required to neutralize 10.0 mL of the acid solution for the equivalent point?

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