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Chemistry 12 - Unit 3

KEY Solubility

1. Write a net ionic equation (balanced, all charges, subscripts, arrows) correct, correct order etc.) for the crystallization of calcium hydroxide. (2 marks)
 $Ca^{2+}(aq) + 2OH^{-}(aq) \rightarrow Ca(OH)_2(s)$

2. Write a net ionic equation (balanced, all charges, subscripts, arrows) correct, correct order etc.) for dissolving ammonium carbonate. (2 marks)
 $(NH_4)_2CO_3(s) \rightarrow 2NH_4^{+}(aq) + CO_3^{2-}(aq)$

3. Write a net ionic equation (balanced, all charges, subscripts, arrows) correct, correct order etc.) for the precipitation of magnesium hydroxide. (2 marks)
 $Mg^{2+}(aq) + 2OH^{-}(aq) \rightarrow Mg(OH)_2(s)$

4. Write a net ionic equation (balanced, all charges, subscripts, arrows) correct, correct order etc.) for the equilibrium between a saturated solution of calcium oxalate. (2 marks)
 $CaC_2O_4(s) \rightleftharpoons Ca^{2+}(aq) + C_2O_4^{2-}(aq)$

11. Calcium fluoride has a solubility of 6.87 grams/L at a certain temperature. Express this solubility in moles per litre. (2 marks) (Show all work. Include units in your answer. Use correct sig. figs.)
 $6.87 \frac{g}{L} \times \frac{1 \text{ mol}}{78.1g} = 0.0880 \text{ mol/L} \quad (8.80 \times 10^{-2} \text{ mol/L})$
 Answer: 0.0880 mol/L

12. The molar solubility of Ag_2CO_3 at a certain temperature is $8.3 \times 10^{-5} \text{ M}$. Express this solubility in grams per litre. (2 marks) (Show all work. Include units in your answer. Use correct sig. figs.)
 $8.3 \times 10^{-5} \frac{\text{mol}}{L} \times \frac{275.8g}{1 \text{ mol}} = 0.023g/L \quad (2.3 \times 10^{-2} g/L)$
 Answer: $0.023g/L$

13. 0.0021 grams of $MgCO_3$ will dissolve in 1.0 L of water at a certain temperature. Express this solubility in grams/100 mL of water. (2 marks) (Show all work. Include units in your answer. Use correct sig. figs.)
 $\frac{0.0021g}{1.0L} \times \frac{100mL}{1L} = 0.00021g/100mL \text{ H}_2O$
 Answer: $(2.1 \times 10^{-4} g/100mL \text{ H}_2O)$

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