

# Download File PDF Shivaji University Kolhapur Sample Question Papers Engineering Be First Year

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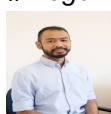
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F. E. - Sem II (C&G) 1/10/15  
App. Chemistry - II  
Q.P. Code : 5816  
(2 Hours) | Total Marks : 60

N.B. : (1) Question No. 1 is compulsory.  
(2) Attempt any three questions out of the remaining five questions.  
(3) All questions carry equal marks.  
(4) Figures to the right indicate full marks.  
(5) Atomic weights: H=1, C=12, N=14, O=16, S=32, Cl=35.5

1. Answer any five of the following - 15

- What is 'Oxidation corrosion'? Why do gold and platinum metal get corroded in atmospheric oxygen?
- Give composition, properties and uses of German silver.
- What is 'cracking of heavy oil'? Mention any four advantages of catalytic cracking over thermal cracking.
- Explain prevention of waste principle in Green Chemistry.
- What are composite materials? Mention any four characteristics of composite materials.
- What is metal cladding? How is nickel obtained?

(a) 1.5g of a coal sample was burnt in a combustion apparatus and the products of combustion were collected in previously weighed KOH bulb and CaCl<sub>2</sub> tube. The increase in weights of KOH bulb and CaCl<sub>2</sub> tube were found to be 3.92g and 1.23g respectively. Calculate percentage carbon and hydrogen in the sample.

2. (a) How do the following factors affect the rate of corrosion? 6  
(i) Relative area of anodic and cathodic parts  
(ii) Purity of metal  
(iii) pH of medium.

(b) What is meant by 'Knocking' in internal combustion engine? Define Octane number. Name any two anti-knock agents. 5

(c) Calculate percentage atom economy for the following reaction with respect to cinnamaldolaldehyde. 4

$$\text{C}_6\text{H}_5\text{CHO} \xrightarrow{\text{O}_2} \text{C}_6\text{H}_5\text{COOH} \xrightarrow{\text{NaOH}} \text{C}_6\text{H}_5\text{COONa} + \text{H}_2\text{O}$$

benzaldehyde acetaldehyde cinnamaldolaldehyde

3. (a) A gaseous fuel has the following composition by volume. 6  
CO = 40% H<sub>2</sub> = 30% CH<sub>4</sub> = 10%  
O<sub>2</sub> = 4% N<sub>2</sub> = 1% CO<sub>2</sub> = 2% and O<sub>2</sub> = 7%  
Calculate volume and weight of air required for complete combustion of 1m<sup>3</sup> of fuel (Mol. wt. of air = 28.947) [TURN OVER

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