

# Download File PDF Review Sheet Unit 6 Chemistry Answer Key

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## Chemistry – Unit 6 Worksheet 2

Balance the following equations by inserting the proper coefficients.

1.  $\text{SO}_2 + \text{O}_2 \rightarrow \text{SO}_3$
  2.  $\text{CH}_4 + \text{O}_2 \rightarrow \text{CO} + \text{H}_2\text{O}$
  3.  $\text{P} + \text{O}_2 \rightarrow \text{PO}_3$
  4.  $\text{CO} + \text{O}_2 \rightarrow \text{CO}_2$
  5.  $\text{CH}_4 + \text{O}_2 \rightarrow \text{CH}_3\text{OH}$
  6.  $\text{Li} + \text{Br}_2 \rightarrow \text{LiBr}$
  7.  $\text{Al}_2\text{O}_3 \rightarrow \text{Al} + \text{O}_2$
  8.  $\text{Na} + \text{H}_2\text{O} \rightarrow \text{NaOH} + \text{H}_2$
  9.  $\text{CO}_2 + \text{H}_2\text{O} \rightarrow \text{C}_6\text{H}_{12}\text{O}_6 + \text{O}_2$
  10.  $\text{H}_2\text{SO}_4 + \text{NaCl} \rightarrow \text{HCl} + \text{Na}_2\text{SO}_4$
- Two reactions used to get rid of sulfur dioxide, a pollutant from burning coal:
11.  $\text{H}_2 + \text{SO}_2 \rightarrow \text{H}_2\text{S} + \text{H}_2\text{O}$
  12.  $\text{CaCO}_3 + \text{SO}_2 + \text{O}_2 \rightarrow \text{CaSO}_4 + \text{CO}_2$
13.  $\text{AgNO}_3 + \text{CaCl}_2 \rightarrow \text{AgCl} + \text{Ca(NO}_3)_2$
  14.  $\text{HCl} + \text{Ba(OH)}_2 \rightarrow \text{BaCl}_2 + \text{H}_2\text{O}$
  15.  $\text{H}_3\text{PO}_4 + \text{NaOH} \rightarrow \text{Na}_3\text{PO}_4 + \text{H}_2\text{O}$
  16.  $\text{Pb(NO}_3)_2 + \text{KI} \rightarrow \text{PbI}_2 + \text{KNO}_3$
  17.  $\text{CuO} + \text{NH}_3 \rightarrow \text{N}_2 + \text{Cu} + \text{H}_2\text{O}$
  18.  $\text{C}_2\text{H}_5\text{OH} + \text{O}_2 \rightarrow \text{CO}_2 + \text{H}_2\text{O}$
  19.  $\text{C}_2\text{H}_6 + \text{O}_2 \rightarrow \text{CH}_3\text{COOH} + \text{H}_2\text{O}$
  20.  $\text{NO}_2 + \text{H}_2\text{O} \rightarrow \text{HNO}_3 + \text{NO}$

Write the formulas of the reactants and products - including the symbols for the state - (s), (l), (g), (aq) - then balance the equations.  
1. Gaseous ammonia ( $\text{NH}_3$ ) reacts with hydrogen chloride gas to form solid ammonium chloride.

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