

Download File PDF Reversible Reactions And Equilibrium Workbook Answers

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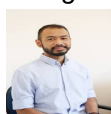
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Reversible Reactions And Equilibrium Workbook Answers

Section 18.2
Reversible Reactions and Equilibrium

OBJECTIVES

- Describe how the amounts of reactants and products change in a chemical system at equilibrium.
- Identify three stresses that can change the equilibrium position of a chemical system.
- Explain what the value of K_{eq} indicates about the position of equilibrium.

Reversible Reactions

- Some reactions do not go to completion as we have assumed. They may be _____ a reaction in which the conversion of reactants to products and the conversion of products to reactants occur simultaneously.
 $2SO_2(g) + O_2(g) \rightleftharpoons 2SO_3(g)$
- The two equations can be combined into one, by using a _____, which tells us that it is a reversible reaction:
 $2SO_2(g) + O_2(g) \rightleftharpoons 2SO_3(g)$
- A chemical equilibrium occurs, and _____ change occurs in the actual amounts of the components of the system.
- Even though the rates of the forward and reverse are equal, the _____ of components on both sides may not be equal.
- An equilibrium position may be shown:
A: _____ or A: _____ B: _____
1% 99% 99% 1%
- Note the emphasis of the arrow direction.
- It depends on which side is favored; almost all reactions are reversible to some extent.

Le Chatelier's Principle

- The French chemist Henri Le Chatelier (1850-1936) studied how the equilibrium position shifts as a result of changing conditions.
- Le Chatelier's principle: _____

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