

# Download File PDF Reaction And Equilibrium Workbook Answers

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Chemistry 12 Unit 2 - Chemical Equilibrium **KEY**

26. A system has reached equilibrium when:

- maximum entropy has been achieved
- minimum entropy has been achieved
- the rate of the forward reaction and reverse reaction is zero
- the concentrations of reactants and products have stopped changing

Your answer is d. Explain why. Again, conc is a macroscopic property - stops changing once equilibrium is reached

27. Equilibrium is achieved when reactant and product concentrations are (equilibrium) constant

28. In a particular chemical reaction,  $\Delta H = +100 \text{ kJ}$ . When equilibrium has been established, it is found that a significant amount of product has formed, even though there is still some reactant left.

What has happened to entropy as this reaction was taking place? increased

Explain how you arrived at your answer. Ex. is endothermic (absorbs) so tendency to min entropy favors reactants. In order for this to occur at all, the tend to max entropy must favor products (products have greater entropy)

29. Given the following potential energy diagram for a reaction:

Explain in terms of enthalpy and entropy, how you could end up with a fairly high ratio of products to reactants. Ex. is all the reactants tend to min entropy, which favors reactants. If reactants ratio is high, tend to max entropy, must favor products (products must have higher entropy)

Worksheet 2-1 - Equilibrium, Enthalpy and Entropy Page 5

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