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so many fake sites. this is the first one which worked! Many thanks

Solution Stoichiometry
Easy Tutorial

5.0 L $\text{AlCl}_3(\text{aq}) \times \frac{0.30 \text{ mol AlCl}_3}{1 \text{ L}} \times \frac{3 \text{ NaOH}}{2 \text{ AlCl}_3}$

EXAMPLE 2:
For the reaction $2\text{Al}(s) + 3\text{FeCl}_2(\text{aq}) \rightarrow 2\text{AlCl}_3(\text{aq}) + 3\text{Fe}(s)$ (is it balanced)
What volume of 0.45 M FeCl_2 is needed to react with 35.0 g Al?

$35.0 \text{ g Al} \times \frac{1 \text{ mol}}{27.0 \text{ g Al}} \times \frac{3 \text{ FeCl}_2}{2 \text{ Al}} \times \frac{1 \text{ L}}{0.45 \text{ mol FeCl}_2} =$

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