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Cool! I'am really happy

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so many fake sites. this is the first one which worked! Many thanks

CHEMISTRY IX NOTES SOLUTION & SUSPENSION – THEORY & QUESTION ANSWERS

MOLE FRACTION (X):
Definition:
Mole fraction (X) of any component in a solution is the number of moles of the component divided by total number of moles making up a solution. It is denoted by X.

Formula:
$$X_{\text{solute}} = \frac{\text{No. of moles of solute}}{\text{Total moles of solute and solvent in the solution}}$$

Example:
For example, a solution is prepared by dissolving 1 mole of ethyl alcohol $\text{C}_2\text{H}_5\text{OH}$ in 3 moles of water (H_2O), where n_1 and n_2 represent the number of moles of ethyl alcohol and water respectively.

Mole fraction of ethyl alcohol = $X_A = \frac{n_1}{n_1 + n_2} = \frac{1}{1 + 3} = \frac{1}{4} = 0.25$

Mole fraction of water = $X_B = \frac{n_2}{n_1 + n_2} = \frac{3}{1 + 3} = \frac{3}{4} = 0.75$

Result:
Mole fraction of ethyl alcohol = $X_A = 0.25$
Mole fraction of water = $X_B = 0.75$

Note, that sum of the mole fractions is equal to 1.
Mole fraction of ethyl alcohol = 0.25
Mole fraction of water = 0.75
Sum of the mole fractions = 1.00

REMEMBER:
The mole fraction is dimensionless quantity that expresses the ratio of the number of moles of one component to the number of moles of all components present. The sum of mole fractions of all components of a solution must equal 1.

PERCENTAGE (%):
Percent concentration is based on mass (M) and volume (V) of the components solute and solvent in the solution. There are four different ways in which percentage concentration can be expressed.

➤ Percentage in $\frac{m}{m}$ = $\frac{\text{Mass } \%}{\text{Mass } \%}$

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